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Information:  
Drawing Covalent Compounds For covalent bonding, we often want to draw how the atoms share electrons in the molecule. For example, consider  $\text{CCl}_4$  and  $\text{NF}_3$  as drawn below: Notice that the atoms share electrons so that they all have 8 electrons.

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Answer: + that S has a negative 2 charge. There are two "X" ions to balance out the negative 2 charge of S. Step 2: Criss cross to balance charges. Step 1: Write each ion and each charge. Positive first! Here's what you get: Add parentheses balance charges. Step 2. Criss cross to

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Hint: add your answer to question 1 and four times your answer to question 2. Verify that there are 32 electrons pictured in the drawing of  $\text{CCl}_4$ . Find the sum of all the valence electrons for  $\text{NF}_3$ . (Add how many valence electrons one nitrogen atom has with the valence electrons for three fluorine atoms.)

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